I'm not robot!

ADVANCED PLACEMENT CHEMISTRY EQUATIONS AND CONSTANTS Throughout the test the following symbols have the definitions specified unless otherwise noted.

 $L_s mL = liter(s), milliliter(s)$ = gram(s) nm = nanometer(s)

= atmosphere(s)

mm Hg = millimeters of mercury J, kJ = joule(s), kilojoulc(s) = volt(s) = mole(s)

ATOMIC STRUCTURE

 $E = h \nu$ $c = \lambda \nu$ E = energy $\nu =$ frequency λ = wavelength

Planck's constant, $h = 6.626 \times 10^{-34} \text{ J}_{\odot}$ Speed of light, $c = 2.998 \times 10^8 \, \mathrm{m \, s^{-1}}$ Avogadro's number = 6.022×10^{23} mol⁻¹ Electron charge, $e = -1.602 \times 10^{-19}$ coulomb

EQUILIBRIUM

 $K_c = \frac{|Cf^c|Df^d}{|Af^c|Bf^b}$, where $a A + b B \rightleftharpoons c C + d D$

 $K_p = \frac{(P_C)^f (P_D)^b}{(P_A)^b (P_B)^b}$

 $K_n = \frac{[H^+][\Lambda^-]}{[H\Lambda]}$

 $K_{p} = \frac{[OH^{-}][HB^{+}]}{[B]}$

 $K_w = [H^*][OH^-] = 1.0 \times 10^{-18} \text{ at } 25^{\circ}\text{C}$

 $= K_a \times K_b$

 $pH = -log[H^*], pOH = -log[OH^*]$

14 = pH + pOH

 $pH = pK_o + log \frac{|A^-|}{|HA|}$

 $pK_a = -\log K_a$, $pK_b = -\log K_b$

 $I_{1/2} = \frac{0.693}{k}$

Equilibrium Constants

Ke (molar concentrations)

K, (gas pressures)

K, (weak acid)

K_k (weak base)

 K_{w} (water)

KINETICS

 $\ln(A)_{i} - \ln(A)_{i} = -kt$

k = rate constantI = time $t_{1/2}$ = half-life

12 AP Chemistry Practice Exam

AP Biology Reference Tables





8. Unpolarised light is incident from air on a plane 11. The ratio of kinetic energy to the total energy of surface of a material of refractive index '\u03c4'. At a particular angle of incidence "i, it is found that the reflected and refracted rays are perpendicular to each other. Which of the following options is correct for this situation?

(1) $i = \sin^{-1}\left(\frac{1}{\mu}\right)$

(2) Reflected light is polarised with its electric vector perpendicular to the plane of incidence

(3) Reflected light is polarised with its electric vector parallel to the plane of incidence

(4) $\mathbf{i} = \tan^{-1} \left(\frac{1}{\mu} \right)$

9. In Young's double slit experiment the separation d between the slits is 2 mm, the wavelength \(\lambda \) of the light used is 5896 A and distance 1) between the screen and slits is 100 cm, It is found that the angular width of the tringes is 0-20°. To increase 13. the fringe angular width to 0.21° (with same & and D) the separation between the slits needs to

(1) 2·1 mm (2) 1.9 mm

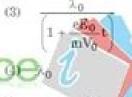
be changed to

- (3) 1·8 mm
- (4) 1·7 mm

10. An astronomical refracting telescope will have large angular magnification and high angular resolution, when it has an objective lens of

- (1) large focal length and large diameter
- (2) large focal length and small diameter (3) small focal length and large diameter
- (4) small focal length and small diameter

- an electron in a Bohr orbit of the hydrogen atom,
- (1) 2:-1
- (2) 1:-1
- (3) 1:1 (4) 1:-2
- 12. An electron of mass m with an initial velocity $\overrightarrow{V} = V_0 \hat{i} \ (V_0 > 0)$ enters an electric field $\vec{E} = -\vec{E_0} \hat{i}$ ($\vec{E_0} = \text{constant} > 0$) at t = 0. If λ_0 is its de-Broglie wavelength initially, then its de-Broglie wavelength at time t is
 - (1) $\lambda_0 t$



radioactive material, half-life is 10 minutes. If initially there are 600 number of nuclei, the time taken (in minutes) for the disintegration of 450 nuclei is

- (1) 30 (2) 10
- (3) 20

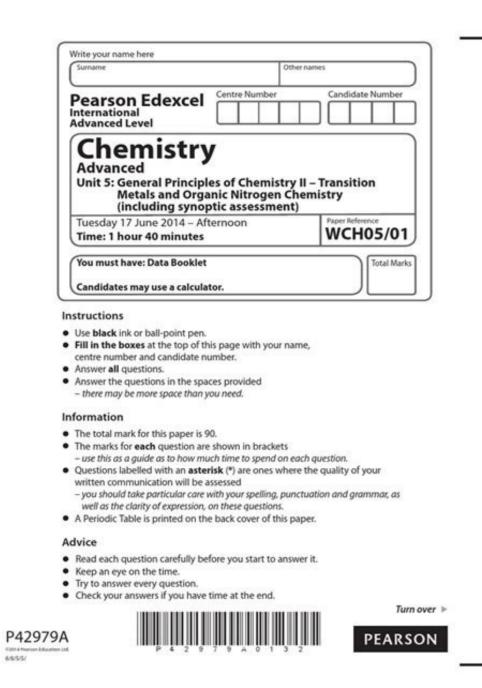
When the light of frequency 2v0 (where vo is threshold frequency), is incident on a metal plate, the maximum velocity of electrons emitted is v1. When the frequency of the incident radiation is increased to 5vo the maximum velocity of electrons emitted from the same plate is v₂. The ratio of v₁ to v₂ is

- (1) 4:1 (2) 1:4
- (3) 1:2(4) 2:1

ACHLA/AA/Page 3

SPACE FOR ROUGH WORK

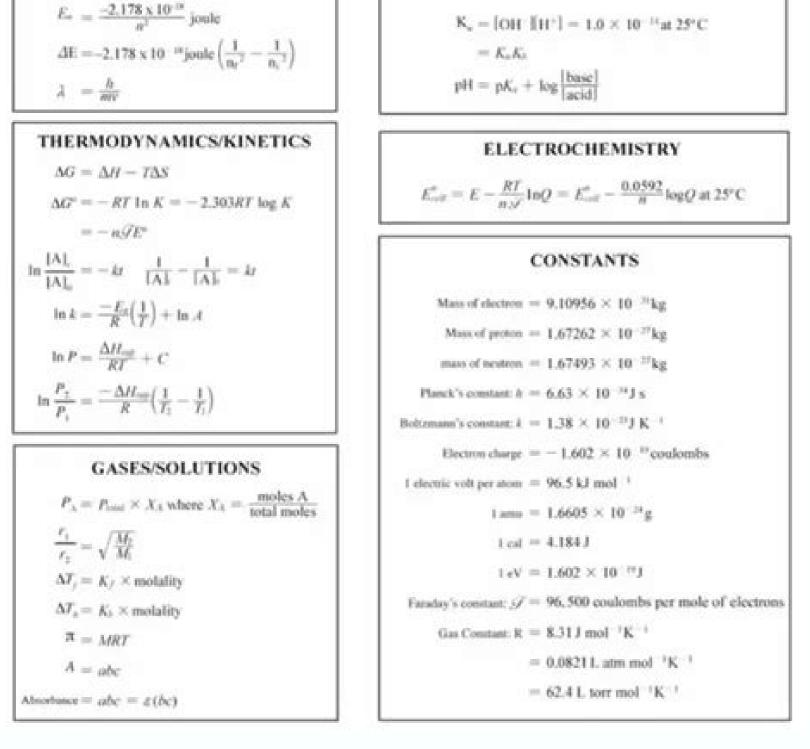
English



ATOMIC STRUCTURE

Chemistry Formulas and Constants

EQUILIBRIUM



Chemistryselect reference style.

Copyright 2001-2018 Castle Software, Inc. All Rights Reserved Students in AP Chemistry need to memorize a lot of information, but we have some good news! During the AP exam, you'll have access to a formula sheet that includes a periodic table as well as useful equations. However, you'll need to have a good sense of what's on the AP Chem

formula sheet and how to use it in order to make the most of it during the test. In this guide, we explain everything you need to know about the AP Chemistry reference table. We go over what the formula sheet looks like, what it includes, which it is included, which is inclu What's on the AP Chem Formula Sheet? The AP Chemistry reference sheet is actually three pages of information. It will be at the beginning of both the multiple-choice and free-response sections of the test, so you'll have access to it for the entire AP Chemistry exam. This means you don't need to memorize any of the information the formula sheet contains. So what's on this AP Chem reference sheet? Check it out here on pages 2-4, as well as the screenshots below (click on each image to enlarge). The AP Chem equation sheet covers six main topics. Periodic Table of Elements The AP Chem equation sheet includes a complete periodic table. For each element, it includes the element, and atomic number, and atomic number number, and atomic number, and atomic number, and atomic number number, and atomic number number, and atomic number number, and atomic number nu Gas constant (3) 1 atm STP Ideal gas at STP Thermodynamics/Electrochemistry 9 equations Values for: What Isn't on the AP Chem Formula Sheet That You Need to Know for the exam, there is some information it's missing. That's what we cover in this section. Be sure to know these equations for the exam. Rate Law One of the laws you learn in AP Chemistry is that, at a constant temperature, the rate of a chemical reaction depends only on the concentrations of the substances that influence the rate. These substances are typically reactants, but can also include products and catalysts. The rate law formula is: rate=k[A]a[B]b [A] and [B] represent molar concentrations of reaction and the temperature) a and b are typically positive integers that must be found experimentally Rate law is used to estimate the relationship between the rate of a reaction and the concentrations of reactants. You may see it on questions related to kinetics. Coulomb's Law Coulomb's Law States that magnitude of the charges and inversely proportional to the square of the distance between them. You may see it on electrochemistry questions. F=k(q1q2/r) r= distance of separation Important Percentages You might need to use these equations on any portion of the test. They determine how accurate estimates are compared to actual results. %error = ((actual value)/(theoretical value))*100 %yield = (actual yield/theoretical yield)*100 Looking for help studying for your AP exam? Our one-on-one online AP tutoring services can help you prepare for your AP exams. Get matched with a top tutor who got a high score on the exam you're studying for! 4 Tips for Getting the Most Out of the AP Chem Equation Sheet The properly. Here are four tips for getting the most out of it. #1: Know How to Use Every Equation on the AP Chem Formula Sheets, so if a formula is included on the sheet, that means there's a solid chance you'll need to use it on the exam. You don't want to waste your exam trying to learn how to use, say Planck's equation, and hoping you got it right! Well before exam day, be sure to go through every formula on the sheet and make sure you understand it and know how to use it. The keys included in the formula sheet are also helpful for figuring out what different variables represent in the equations, so don't neglect them either. #2: Use You're allowed a graphing calculator by programming constants and formulas into it so you can solve equations more easily. However, don't think this means you get out of understanding the actual formulas themselves. You still need to know when and how to use each equation on the AP Chemistry reference table, not to mention you won't have your calculator for multiple-choice questions. So know when your calculator is useful, but don't over-rely on it. #3: Take Practice Tests Using the AP Chem formula sheet is key to doing well on the exam. You should take several AP Chem formula sheet is key to doing well on the exam. For each of them, use the official formula sheet. Your teacher will also likely give you a copy of the AP Chem formula sheet for your in-class exams, so you can get some practice in there too. If you need help finding practice tests, check out our guide specifically on where to find the best AP Chem formula sheet for your in-class exams, so you can get some practice in there too. If you need help finding practice tests, check out our guide specifically on where to find the best AP Chem formula sheet for your in-class exams, so you can get some practice in there too. If you need help finding practice tests, check out our guide specifically on where to find the best AP Chem formula sheet for your in-class exams, so you can get some practice in there too. If you need help finding practice tests, check out our guide specifically on where to find the best AP Chem formula sheet for your in-class exams, so you can get some practice in the practice tests, check out our guide specifically on where to find the best AP Chem formula sheet for your in-class exams, so you can get some practice in the pract the AP Chemistry exam, it's that you'll need to look at the included periodic table during the test. The periodic table included in the AP Chemistry equation sheet has all the information you'll need to answer questions about it. However, that doesn't mean you can ignore it until it comes time to take the test. It's very important to be familiar with the periodic table throughout the year. You should know what each value in the table represents, how to use group numbers to determine the number of bonds and valence electrons an element can form, and what the key periodic trends are. Spend the most time on groups 1-8A, as you'll be tested on these elements most often. What's Next? Learn the difference between Physics 1, Physics 2, and Physics 1, Physics 2, and Physics 1, Physics 2, and Physics 2 and Physics 2 and Physics 2. top 5 strategies you must be using to have a shot at improving your score. Download it for free now: Copyright 2001-2018 Castle Software, Inc. All Rights Reserved

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